## **UNIT TEST-04**

Subject : Chemistry Class : XI

| <b>Q.1</b> (1)  | <b>Q.2</b> (2)  | <b>Q.3</b> (1)  | <b>Q.4</b> (3)  | <b>Q.5</b> (2)  | <b>Q.6</b> (3)  | <b>Q.7</b> (3)  | <b>Q.8</b> (3)  | <b>Q.9</b> (1)  | Q.10(2)         |
|-----------------|-----------------|-----------------|-----------------|-----------------|-----------------|-----------------|-----------------|-----------------|-----------------|
| <b>Q.11</b> (2) | Q.12(2)         | <b>Q.13</b> (1) | <b>Q.14</b> (3) | <b>Q.15</b> (2) | <b>Q.16</b> (4) | <b>Q.17</b> (2) | <b>Q.18</b> (3) | <b>Q.19</b> (4) | <b>Q.20</b> (3) |
| <b>Q.21</b> (2) | <b>Q.22</b> (2) | <b>Q.23</b> (2) | <b>Q.24</b> (2) | <b>Q.25</b> (2) | <b>Q.26</b> (2) | <b>Q.27</b> (1) | Q.28(2)         | <b>Q.29</b> (4) | <b>Q.30</b> (3) |
| <b>Q.31</b> (3) | <b>Q.32</b> (4) | <b>Q.33</b> (3) | <b>Q.34</b> (3) | <b>Q.35</b> (4) | <b>Q.36</b> (3) | <b>Q.37</b> (4) | <b>Q.38</b> (2) | <b>Q.39</b> (2) | <b>Q.40</b> (4) |
| <b>Q.41</b> (3) | <b>Q.42</b> (3) | <b>Q.43</b> (2) | <b>Q.44</b> (1) | <b>Q.45</b> (3) | <b>Q.46</b> (4) | <b>Q.47</b> (2) | <b>Q.48</b> (4) | <b>Q.49</b> (3) | <b>Q.50</b> (4) |

Q.1 (1)  $CH_2 = CH - CH_2 - C \equiv N$   $\sigma - bond$   $\pi - bond$  9 3Ratio 3: 1

Q.2 (2) Stability order

$$\overbrace{\bigcirc \\ NO_2}^{\overline{C}H_2} > \overbrace{\bigcirc \\ F}^{\overline{C}H_2} > \overbrace{\bigcirc \\ CH_2}^{\overline{C}H_2} > \overbrace{\bigcirc \\ CH_2}^{\overline{C}H_2}$$

presence of [-I] group will increase the stability of carbanion.

**Q.3** (1)

**Q.4** (3)

-Cl shows +m effect that's why it is O, P-directing

**Q.5** (2)

**Q.6** (3)

$$COOH$$
  $COOH$   $COOH$   $OH$   $COOH$   $CO$ 

Q.7 (3) Ortho effect.

**Q.8** (3)

**Q.9** (1)

Naphthalene

$$(14\pi \text{ electrons})$$

**Q.10** (2)

**Q.11** (2)

Q.12 (2)

**Q.13** (1)

Q.14 (3)
In the compound 3 (C = C) with different substituent so total G. Isomer = 6

**Q.15** (2)

**Q.16** (4)

Q.17 (2)  ${}^{4}_{CH_{3}} - {}^{3}_{CH} = {}^{2}_{CH} - {}^{1}_{CN}$  ${}^{8}_{p^{3}} + {}^{3}_{p^{2}} + {}^{2}_{p^{2}} + {}^{3}_{p^{2}}$ 

**Q.18** (3)

n it is not homocyclic since hetero atom is present H
within the ring.

**Q.19** (4)

**Q.20** (3)

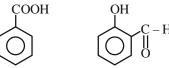
**Q.21** (2)

$$\begin{array}{ccc}
CH_3 - CH - CH_2 - OH & 2 - Bromopropanol \\
& Br
\end{array}$$

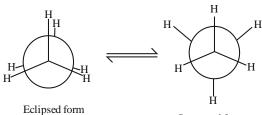
Q.22 (2)

- Q.23 (2)
- Q.24 (2)
- Q.25 (2) Best for symmetrical alkane.
- Q.26 (2) CHO Formed only one ozonized product that is
- Q.27 (1) Hydrogenation  $\rightarrow$  H<sub>2</sub>/Ni De HydroHologenation  $\rightarrow$  Ale · KOH De Hydration  $\rightarrow$  Conc  $\cdot$  H<sub>2</sub>SO<sub>4</sub>/ $\Delta$ Hydration  $\rightarrow H_3O^{\oplus}$
- Q.28 (2)
- Q.29 (4)
- Q.30 (3)
- Q.31 (3)
- Q.32 (4) Acidic strength ∞ Stability of conjugate base E.N.  $\rightarrow$  sp carbon > sp<sup>2</sup> carbon > sp<sup>3</sup> carbon
- Q.33 (3) CH<sub>3</sub> Conc HNO Sn/HCl NO,

- Q.34 (3) Hint:- decolourise bromine water do not show properties of Aromatic compounds
- Q.35 Reaction is called Wurtz fittig reaction.
- Q.36 (3) $CH_2 - CH - CH_3$ 
  - 7, 2'H are present in the compound.
- Q.37 All are correct statement (fact Based)
- Q.38 (2) COOH



- Both are functional isomer of each other.
- Q.39 (2)
- Q.40  $CH_3 - CH_3 \Rightarrow Ethane$



Staggered form

- both are inter conver Rapidly due to free rotation so it can not be isolated at room temperature.
- Q.41 (3)
- Q.42 IUPAC name – Chlorophenyl methane
- Q.43 (2)
- 0.44 (1)
- Q.45 (3)
- (4) Q.46

Q.47 (2)

(2) Acidic strength of hydrogen

$$CH = CH > CH_2 = CH_2 > CH_3 - CH_3$$

**Q.48** (4)

$$CH_{3}-C=C-CH_{2}-CH_{3}$$

$$CH_{3}-C=C$$

$$H_{2}/Pd$$

$$BaSO_{4}$$

$$(A)$$

$$Na/liq.NH_{3}$$

$$CH_{3}$$

$$CH_{3}$$

$$CH_{3}$$

$$CH_{2}$$

$$CH_{2}CH_{3}$$

$$(B)$$

**Q.49** (3)

**Q.50** (4)